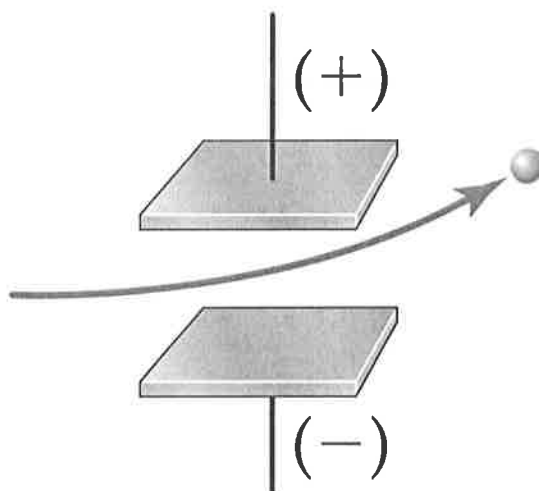


Ions Charges-- Practice #1

Name: \_\_\_\_\_

Using a periodic table, fill in the following table.

	Number of (+)Protons	Number of (-) Electrons	Overall Charge	Element Symbol & Charge	Neutral (N) Cation (C) Anion (A)
1.	+7	-10			
2.	+12	-10			
3.	+30	-30			
4.	+19		1+		
5.				P <sup>3-</sup>	
6.	+38		2+		
7.	+8		0		
8.				Cr <sup>3+</sup>	
9.	+35	-36			
10.	+55		1+		



11. What is the charge on the particle traveling between the two charged plates?

(Hint: Think about magnets)

12. How do you know?

13. Would the particle most likely be a proton or electron? Why?

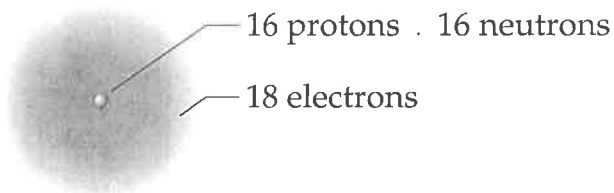
## Ion Class Sheet #2

Name: \_\_\_\_\_

**Ion** – An atom with an overall charge due to unequal numbers of protons (+) and electrons (-).

**Anion** – Ion with an overall negative (-) charge.

**Cation** – Ion with an overall positive (+) charge.



1. Try This: (Use the picture above)

Name of where the protons and neutrons live:

Name of where the electrons live:

Atomic Number:

Number of protons:

Number of neutrons:

Atomic Mass:

Number of electrons:

Overall Charge:

Symbol with charge:

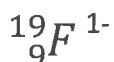
2. Try This:

	Number of (+) Protons	Number of (-) Electrons	Overall Charge	Element Symbol & Charge	Neutral (N) Cation (C) Anion (A)
1.	+9	-10	-1	F <sup>1-</sup>	A
2.	+20	-18			
3.	+7	-7			
4.	+3		1+		
5.				Al <sup>3+</sup>	
6.	+19	-18			
7.				S	
8.				I <sup>1-</sup>	
9.	+26	-23			
10.				Ba <sup>2+</sup>	

9PS - Ion Notation Worksheet 1

Name: \_\_\_\_\_

1. Given the information and a periodic table - complete the following.



element: \_\_\_\_\_

# protons: \_\_\_\_\_

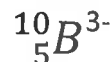
# neutrons: \_\_\_\_\_



element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# neutrons: \_\_\_\_\_



mass #: \_\_\_\_\_

# protons: \_\_\_\_\_

# electrons: \_\_\_\_\_

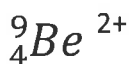


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

# electrons: \_\_\_\_\_

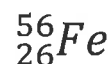


element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# electrons: \_\_\_\_\_

atomic mass: \_\_\_\_\_

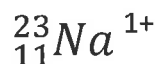


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

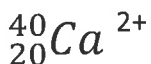


element: \_\_\_\_\_

atomic #: \_\_\_\_\_

mass #: \_\_\_\_\_

# electrons: \_\_\_\_\_

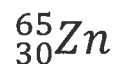


element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

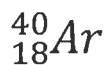


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

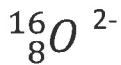


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

# electrons: \_\_\_\_\_



element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# electrons: \_\_\_\_\_

atomic mass: \_\_\_\_\_



element: \_\_\_\_\_

# protons: \_\_\_\_\_

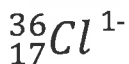
# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

9PS - Ion Notation Worksheet 2

Name: \_\_\_\_\_

1. Given the information and a periodic table - complete the following.

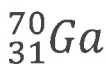


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

# electrons: \_\_\_\_\_



element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# neutrons: \_\_\_\_\_

# electrons: \_\_\_\_\_

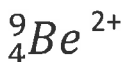


mass #: \_\_\_\_\_

# protons: \_\_\_\_\_

# electrons: \_\_\_\_\_

atomic #: \_\_\_\_\_

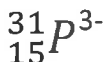


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

# electrons: \_\_\_\_\_

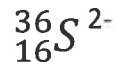


element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# electrons: \_\_\_\_\_

atomic mass: \_\_\_\_\_

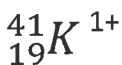


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

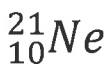


element: \_\_\_\_\_

atomic #: \_\_\_\_\_

mass #: \_\_\_\_\_

# electrons: \_\_\_\_\_

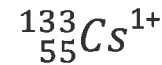


element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_



element: \_\_\_\_\_

# protons: \_\_\_\_\_

# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

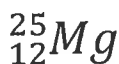


element: \_\_\_\_\_

# protons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

# electrons: \_\_\_\_\_

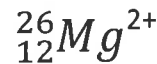


element: \_\_\_\_\_

atomic #: \_\_\_\_\_

# electrons: \_\_\_\_\_

atomic mass: \_\_\_\_\_



mass #: \_\_\_\_\_

# protons: \_\_\_\_\_

# electrons: \_\_\_\_\_

# neutrons: \_\_\_\_\_

What is the term used to describe these last examples? \_\_\_\_\_

What do all three have in common? \_\_\_\_\_