WRITING BINARY FORMULAS

Name _____

Write the formulas for the compounds formed from the following ions.

1. Na^{+}, Cl^{-} _____
2. Ba^{+2}, F^{-} _____
3. K^{+}, S^{-2} _____
4. Li^{+}, Br^{-} _____
5. Al^{+3}, I^{-} _____
6. Zn^{+2}, S^{-2} _____
7. Ag^{+}, O^{-2} _____
8. Mg^{+2}, P^{-3} _____
9. Ni^{+2}, O^{-2} _____
10. Ni^{+3}, O^{-2} _____
11. Fe^{+2}, O^{-2} _____
12. Fe^{+3}, O^{-2} _____
13. Cr^{+2}, S^{-2} _____
14. Cr^{+3}, S^{-2} _____
15. Cu^{+}, Cl^{-} _____
16. Cu^{+2}, Cl^{-} _____
17. Pb^{+2}, O^{-2} _____
18. Pb^{+4}, O^{-2} _____
19. Mn^{+2}, Br^{-} _____
20. Mn^{+4}, Br^{-} _____

NAMING BINARY COMPOUNDS (IONIC)

Name _____

Name the following ionic compounds using Roman numerals where necessary.

1. $BaCl_2$ _____
2. NaF _____
3. Ag_2O _____
4. CuBr _____
5. $CuBr_2$ _____
6. FeO _____
7. Fe_2O_3 _____
8. MgS _____
9. Al_2O_3 _____
10. CaI_2 _____
11. K_2S _____
12. $CrCl_2$ _____
13. $CrCl_3$ _____
14. CaO _____
15. Ba_3P_2 _____
16. Hg_2I_2 _____
17. Na_2O _____
18. BeS _____
19. MnO _____
20. Mn_2O_3 _____

FORMULAS WITH POLYATOMIC IONS

Name _____

Matching the horizontal and vertical axes, write the formulas of the compounds with the following combination of ions. The first one is done for you.

	OH^-	NO_3^-	CO_3^{-2}	SO_4^{-2}	PO_4^{-3}
H^+	HOH (H_2O)	HNO_3	H_2CO_3	H_2SO_4	H_3PO_4
Na^+					
Mg^{+2}					
NH_4^+					
Ca^{+2}					
K^+					
Al^{+3}					
Pb^{+4}					

NAMING OF NON-BINARY COMPOUNDS

Name _____

An ionic compound that contains more than two elements must contain a polyatomic ion. Name the following compounds.

1. $NaNO_3$ _____
2. $Ca(OH)_2$ _____
3. K_2CO_3 _____
4. NH_4Cl _____
5. $MgSO_4$ _____
6. $AlPO_4$ _____
7. $(NH_4)_2SO_4$ _____
8. Na_3PO_4 _____
9. $CuSO_4$ _____
10. NH_4OH _____
11. Li_2SO_3 _____
12. $Mg(NO_3)_2$ _____
13. $Al(OH)_3$ _____
14. $(NH_4)_3PO_4$ _____
15. KOH _____
16. $Ca(NO_3)_2$ _____
17. K_2SO_4 _____
18. $Pb(OH)_2$ _____
19. Na_2O_2 _____
20. $CuCO_3$ _____