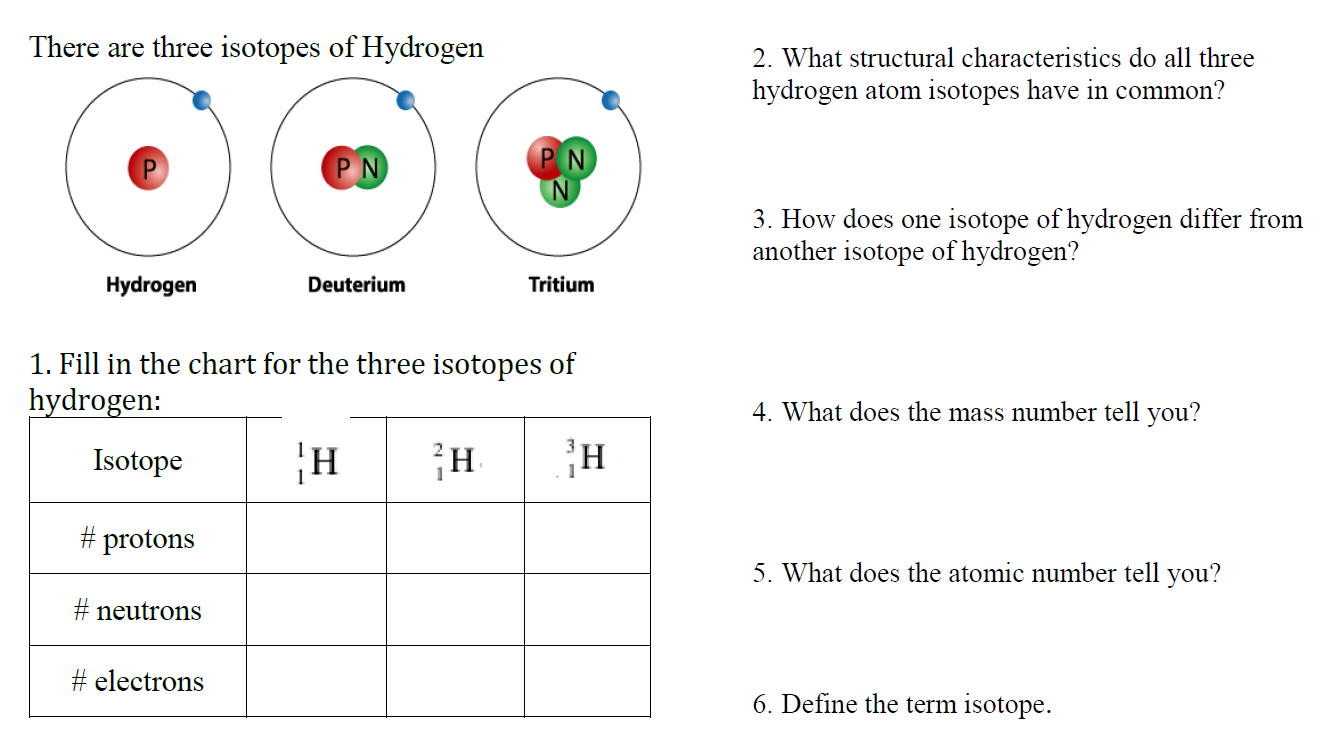
Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period:\_\_\_\_\_\_\_\_

**Isotope Practice**



6

Here are three isotopes of an element: 12C 13C 14C

6

6

* 1. The element is: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
  2. The number 6 refers to the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
  3. The numbers 12, 13, and 14 refer to the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
  4. How many protons and neutrons are in the first isotope? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
  5. How many protons and neutrons are in the second isotope? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
  6. How many protons and neutrons are in the third isotope? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Complete the following chart:

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Isotope name** | **atomic #** | **mass #** | **# of protons** | **# of neutrons** | **# of electrons** |
| Potassium-37 |  |  |  |  |  |
| Oxygen-17 |  |  |  |  |  |
| uranium-235 |  |  |  |  |  |
| uranium-238 |  |  |  |  |  |
| boron-10 |  |  |  |  |  |
| boron-11 |  |  |  |  |  |

